AMENDMENTS TO THE SPECIFICATION

Replace paragraph [0059] with:

[0059] A non-luminescent or secondary colorant (e.g., dye) may be incorporated into the ink composition to provide a visible color to the mark or to alter the luminescence properties of the primary luminescent compound (e.g., dye). Preferred secondary colorants are those that exhibit a color change when the pH of their environment is altered. Even more preferable are those that exhibit negligible absorbance in the near UV (300 to 400 nm wavelength), so as not to interfere with the absorbance of the energy active or photoactive compounds, and that exhibit a high degree of light fastness, e.g., to sunlight. Examples of preferred secondary colorants that exhibit color change under acidic conditions are Basic Violet 4 (triarylmethane dye class, Color Index 42600), Solvent Red 49 (xanthene dye class, Color Index 45170:1), and Solvent Orange 3 (azo dye class, Color Index 11270:1). Other examples of secondary colorants that may be appropriate for the present invention include colorants from the dye classes listed above as well as colorants from dye classes such as anthraquinones, diphenylmethanes, thiazines, oxazines, azines, pyronines, thiopyronines, acridines, polymethine, indigoid, nitro, and nitroso, and combinations thereof. <u>The non-luminescent colorant can be titanium dioxide, a triarylmethane dye, or a xanthene dye.</u>

Replace paragraph [0091] with:

[0091] It is important to ~~not~~ <u>note</u> that in the above example that both the colorant (i.e., Chrysolidine Y base) and the luminescent compound (i.e., LUMILUX CD-331 dye) will simultaneously react with the energy active compound (i.e., tribromoethanol). In fact, both changes serve to diminish the overall luminescence intensity of the luminescent compound. In the case of the luminescent compound, the luminescence is quenched directly. In addition, the change in color of the colorant from yellow to brown corresponds to an increase in the average energy of light absorbed. In the initial state, the yellow dye absorbs primarily in the region from 400-500 nm and is practically transparent in the emission region of the luminescent compound. In the final state, the colorant absorbs light strongly in the same energy region (about 610 nm) that luminescence occurs and thus contributes to the reduction in overall observable luminescence intensity.